

Directions: Highlight each of the terms below. Then, write the correct definition next to each term.

IMPORTANT: Use the provided textbook pages and the links below for definitions. DO NOT GOOGLE THEM.

http://www.chem4kids.com/files/matter_changes.html

http://www.physics4kids.com/files/thermo_expand.html

1. **Phase of Matter**: The states in which matter can exist; including solid, liquid and gas. When temperature changes, matter can undergo a phase change, shifting from one form to another.
2. **Solid**: Phase of matter with a definite shape and a definite volume.
3. **Liquid**: Matter having definite volume but not definite shape.
4. **Gas**: Matter having no definite shape or volume.
5. **Condensation**: When a gas changes into a liquid.
6. **Boiling**: When a liquid turns into a gas.
7. **Evaporation**: When a liquid turns into a gas.
8. **Freezing**: When a liquid turns into a solid.
9. **Sublimation**: When a substance changes directly from a solid to a gas.
10. **Deposition**: When a substance changes directly from a gas to a solid.
11. **Melting**: When a substance changes from a solid to a liquid.
12. **Compress**: When the particles of a substance get closer together.

13. **Expand**: When the particles of a substance spread further apart.
14. **Definite**: Will not change.
15. **Volume**: The amount of space an object takes up.
16. **Heat Energy**: A form of energy produced by vibration of particles that can be absorbed, given up or transferred between substances.
17. **Temperature**: A measurement of amount of heat energy in a substance, determined by the average kinetic energy (movement) of the molecules in the substance.
18. **Pressure**: The force exerted on a surface.
19. **Charles's Law**: The principle stating that the volume of a gas increases as the temperature increases, as long as the pressure remains constant.
20. **Boyle's Law**: The principle stating that the pressure of a gas increases when the volume decreases, as long as temperature remains constant.
21. **G-L Gas Law**: The principle stating that the pressure of a given mass of gas varies directly with the absolute temperature of the gas, when the volume is kept constant.
22. **Absolute Zero**: The temperature at which molecules do not move at all; the lowest temperature matter can have. (This temperature has never been reached.)