Name_			
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Write the symbol for each element in the chemical formula and tell how many atoms of each element are present in the formula.

REMINDERS: *A new capitol letter indicates a new element.

*Subscripts that are OUTSIDE of parentheses get MULTIPLIED by the subscripts that are inside of the parentheses.

Example: $Fe_2(SO_4)_3 = Fe - 2$ atoms,

S - 3 atoms (multiply the 3 outside by the one S inside)

O - 12 atoms (multiply the 3 outside by the 4 O's inside)

1. K₂SO₄

8. Na₂CO₃

2. $C_6H_{12}O_6$

9. AlPO₄

3. $Mg(NO_3)_2$

10. Ba(NO₃)₂

4. Fe(OH)₃

11. Al₂(SO₄)₃

5. Ca₃(PO₄)₂

12. Mg₃(PO₄)₂

6. Fe₂(CO₃)₃

13. AlCl₃

7 Al(OH)₃

14. Cu(NO₃)₂

Write the symbol for each element in the chemical formula and tell how many atoms of each element are present in the formula.

REMINDERS: *A new capitol letter indicates a new element.

*Subscripts that are OUTSIDE of parentheses get MULTIPLIED by the subscripts that are inside of the parentheses.

*COEFFICIENTS, the big numbers in front of the formulas, apply the EVERYTHING that they are in front of and get MULTIPLIED by ALL OF THE OTHER NUMBERS THAT APPLY TO THAT ELEMENT.

Example: 2 Fe₂(SO₄)₃ = Fe - 4 atoms, (multiply the coefficient 2 by the subscript 2)
S - 6 atoms (multiply the coefficient 2 by the 3 outside the parentheses
O - 24 atoms (multiply the coefficient 2 by the 4 O's inside the parenthese and the the 3 outside of the parentheses)

1. 2K₂SO₄

8. 3Na₂CO₃

2. $2C_6H_{12}O_6$

9. 3AlPO₄

3. $3Mg(NO_3)_2$

10. 3Ba(NO₃)₂

4. 2Fe(OH)₃

11. 2Al₂(SO₄)₃

5. 2Ca₃(PO₄)₂

12. $3Mg_3(PO_4)_2$

6. $2\text{Fe}_2(\text{CO}_3)_3$

13. 3AlCl₃

7. 3Al(OH)₃

14. $3Cu(NO_3)_2$